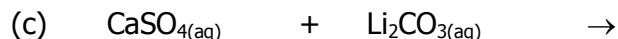
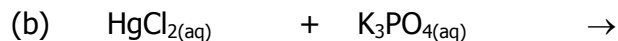
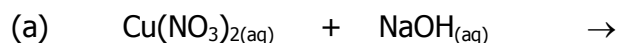


CHEMICAL REACTIONS – SET II

1. Write balanced molecular equations for the following double replacement reactions: (Make sure the products of the molecular equations are electrically neutral.)



2. Write the formulas for the following ionic compounds:

magnesium phosphate _____

sodium sulfide _____

copper(II) iodide _____

3. Identify the cation and anion in each of the following *aqueous* solutions:

$(\text{NH}_4)_2\text{CO}_3$ _____ and _____

CaBr_2 _____ and _____

Ag_2S _____ and _____

4. Using your activity series list the following: three metals that

(a) three metal ions that *can* be replaced by $\text{Zn}(\text{s})$ _____

(b) three metal ions that *cannot* be replaced by $\text{Zn}(\text{s})$ _____

5. When one metal replaces another metal, what is taking place at the atomic level ?

6. List three metal atoms that will lose electrons to the Fe^{2+} ion. _____

7. Write equations for the following single replacement reactions:

